Dear Franc Perdih,

Thanks for your comments on the paper "Synthesis and crystal structure of a polymeric copper(II) complex derived from 2-hydroxy-5-methylbenzaldehyde oxime with antibacterial activities". Please see the revisions below.

Warm regards,

Ya-Li Sang

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Revisions for Reviewer A:

1. Why Authors used out-of-date Sheldrick program? The new version (Sheldrick, G. M. SHELXL-2014/7: Program for the Solution of Crystal Structures, University of Göttingen, Göttingen, Germany, 2014.) should be use as it was released years ago.

Response: Although there is a new version of Shelxl-2014/7, it is of the same accurate to solve and refine structures with Shelxtl-97 version. From my experience, it is not different for the results with different versions of software.

2. Authors claim that the solution of complex in MeCN has non-electrolyte nature with molar conductivity of 25 S/mol for 0.001M solution. This is quite close to the conductivity of LiCl in this solvent. In non-water solvent dissociation is much lower and the value of 25 indicate that dissociation (depolymerization) occurred. How Authors in other way can explain solubility of polymeric copper? In my opinion, the weak (2.571 A) Cu-OH bond breaks and in solution we observe a basic properties of hydroxylamine responsible for high conductivity (pKa = 5.9). The hydroxylaminium ion (from hydrochloride salt) at 25°C has conductivity of 49.27 ±0.17 Ω-1 cm2 g-equiv-1, quite comparable with the value got by Authors. In my opinion Authors should check the variation of molar conductivity with dilution, not only to get limiting conductivity, but also to show, that it is strongly depends on concentration (as for weak electrolytes).

Response: We have determined the conductivity with 0.0001M, which give a value of 7.0Ω–1 cm2 mol–1. In solution, the complex may dissociated to form single [CuL2] units.

3. The UV-VIS spectra. Authors present only bands in the UV region, but Cu(II) complex is blue, indicating on presence of d-d transitions. The relatively high solubility (for conductance measurements 0.001M) should be good enough to see d-d transitions. This could say something about the geometry of copper(II) in solution.

Response: The d-d transition at 605 nm is given, which indicates that the complex is dissociated to form [CuL2] unit with square planar geometry in such solution.

4. Figs. 1 and 2. Authors should write in figure caption the sentence that the hydrogen atoms were omitted for clarity.

Response: Corrected as suggested.

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Revisions for Reviewer B:

Line 9 and 179: should be 2-hydroxy-5-methylbenzaldehyde instead

2-hydroy-5-methylbenzaldehyde

Response: Corrected as suggested.

Table 2. The unit for angle should be ° like it is in the text

Response: Corrected as suggested.

Line 169 the sentence does not correspond to the results from table 3. The right sentence is: The Schiff base HL shows medium antibacterial activities against Escherichia coli and Staphylococcus aureus, and no activity against Bacillus substilis and Pseudomonas fluorescens.

Response: Corrected as suggested.

Line 174-176:

It is not surprising that complex is more active against Escherichia coli and Pseudomonas fluorescens than penicillin. It is well known that penicillin is antibiotic specifically effective against gram-positive bacteria (Staphylococcus aureus Bacillus subtilis) and does not affect most gram-negative bacteria (Escherichia coli and Pseudomonas fluorescens).The other antibiotic standard is required to compare activity of complex against Escherichia coli and Pseudomonas fluorescens. In the description of the results, you should compare antibacterial activity of complex with penicillin only against Staphylococcus aureus and Bacillus subtilis.

Response: Kanamycin G was used as a reference for Escherichia coli and Pseudomonas fluorescens. The discussion has been revised.

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Revisions for Additional Editorial comments:

- Cell parameters in the Abstract are not regarded as informative. Please remove those data.

Response: The cell parameters in the Abstract have been omitted.

- Try to incorporate your own research papers like Acta Chim. Slov. 2016, 63, 856–86 and others.

Response: Corrected as suggested.

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